

# Øving 3

Betrakt følgende irreversibel gassfasereaksjon:



Reaksjonen er utført i en batchreaktor (trykkbeholder) ved 300 °C og 2 atm og med rent A ved start ( $t = 0$ ). Reaksjonen foregikk isotermt og følgende verdier av totaltrykket ( $P_t$ ) ble målt:

| t (min) | $P_t$ (atm) |
|---------|-------------|
| 0       | 2,0         |
| 6,0     | 2,2         |
| 9,8     | 2,3         |
| 14,5    | 2,4         |
| 20,7    | 2,5         |
| 28,5    | 2,6         |
| 40,5    | 2,7         |

- a) Finn et uttrykk for omsetningsgraden ( $X_A$ ) som funksjon av totaltrykket, dvs.  $X_A = f(P_t)$

a) Antar ideell gass, bruker den generelle formelen

$$V = V_0 (1 + \epsilon X) \frac{P_0}{P} \cdot \frac{T_0}{T} \quad \text{Isoterm}$$

$$\Rightarrow (1 + \epsilon X) \frac{P_0}{P} = 1$$

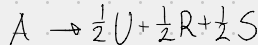
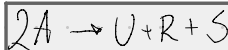
$$\Rightarrow \frac{P_0}{P} + \epsilon X \frac{P_0}{P} = 1$$

$$X = \left(1 - \frac{P_0}{P}\right) \frac{P}{\epsilon P_0}$$

$$X = \frac{P}{\frac{1}{2} \cdot 2} - \frac{1}{\frac{1}{2}}$$

$$\underline{X = P - 2}$$

Siden det er en batchreaktor, er det rimelig å anta  $V = V_0$



$$\Rightarrow \delta = \frac{1}{2} + \frac{1}{2} + \frac{1}{2} - 1 = \frac{1}{2}$$

$$\epsilon = \delta \cdot y_{A_0} = \frac{1}{2}$$

- b) Tegn kurven  $C_A = g(t)$  på vanlig rutepapir

b) Har antatt konstant volum

$$\Rightarrow C_A = C_{A_0} (1 - X) = C_{A_0} (1 - P + 2) = C_{A_0} (3 - P)$$

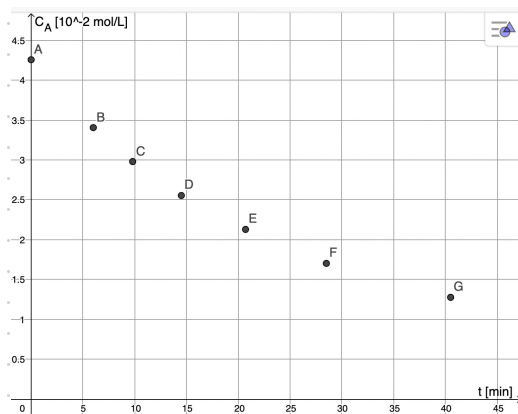
$$\text{Ved ideell gasslov: } C_{A_0} = \frac{N_{A_0}}{V_0} = \frac{P_0}{RT_0} = \frac{2 \text{ atm}}{0,082 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 573 \text{ K}}$$

$$C_{A_0} = 4,257 \cdot 10^{-2} \frac{\text{mol}}{\text{L}}$$

Data punktene blir da:

| t (min) | P_t (atm) | X_A | C_A (mol/L) |
|---------|-----------|-----|-------------|
| 0       | 2         | 0   | 0.04257     |
| 6       | 2.2       | 0.2 | 0.03405     |
| 9.8     | 2.3       | 0.3 | 0.0298      |
| 14.5    | 2.4       | 0.4 | 0.02554     |
| 20.7    | 2.5       | 0.5 | 0.02128     |
| 28.5    | 2.6       | 0.6 | 0.01703     |
| 40.5    | 2.7       | 0.7 | 0.01277     |
| C_A0    | 0.04257   |     |             |

Plotter  $C_A$  mot  $t$ :



c) Finn hastighetskonstanten ( $k$ ) og reaksjonsordenen ( $\alpha$ ) ved 300 °C i hastighetslikningen:

$$-r_A = k \cdot C_A^\alpha$$

g) Designlikningen Batchreaktor:  $\frac{dN_A}{V dt} = r_A$

Ved konstant volum  $V_0$ :  $\frac{1}{V_0} \frac{dN_A}{dt} = r_A$

Ønsker variabelskifte:

$$N_A \rightarrow \frac{N_A}{V_0} \Rightarrow \frac{d(\frac{N_A}{V_0})}{dN_A} = \frac{1}{V_0} \Rightarrow dN_A = V_0 \cdot d(\frac{N_A}{V_0})$$

$$\frac{d(\frac{N_A}{V_0})}{dt} = r_A$$

$$\frac{N_A}{V_0} = C_A$$

$$\frac{dC_A}{dt} = r_A = -k \cdot C_A^\alpha$$

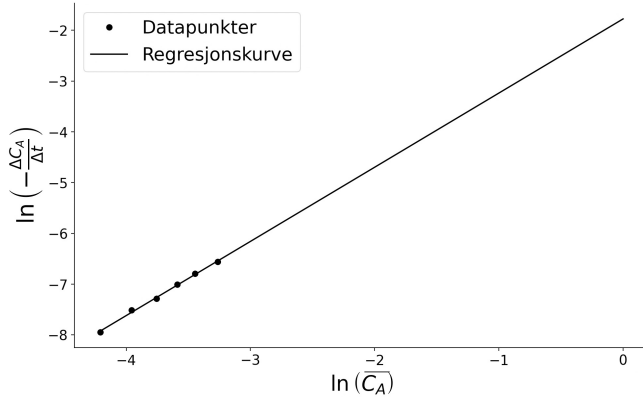
$$-\frac{dC_A}{dt} = k \cdot C_A^\alpha$$

$$\ln\left(-\frac{dC_A}{dt}\right) = \alpha \ln C_A + \ln k$$

Estimerer  $\frac{dC_A}{dt}$  med  $\frac{\Delta C_A}{\Delta t}$ :

Plotter  $\ln\left(\frac{\Delta C_A}{\Delta t}\right) = \alpha \ln C_A + \ln k$  med lineær regresjon:

$$\underbrace{\ln\left(\frac{\Delta C_A}{\Delta t}\right)}_y = \underbrace{\alpha}_{a} \cdot \underbrace{\ln C_A}_x + \underbrace{\ln k}_b$$



For at  $\alpha = 1,461$  og  $\ln k = -1,776 \Rightarrow k = 0,169$

Dermed er  $-r_A = 0,169 C_A^{1,461}$

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import numpy as np
import matplotlib.pyplot as plt
t = np.array([0, 6, 9.8, 14.5, 20.7, 28.5, 40.5])
Pt = np.array([2, 2.2, 2.3, 2.4, 2.5, 2.6, 2.7])
X_A = Pt - 2
C_A0 = 2/(0.082*573)
C_A = C_A0*(3-Pt)
dCA = np.array([C_A[i+1]-C_A[i] for i in range(len(C_A)-1)])
dt = np.array([t[j+1]-t[j] for j in range(len(t)-1)])
snittCA = np.array([(C_A[i+1]+C_A[i])/2 for i in range(len(C_A)-1)])

x = np.log(snittCA)
y = np.log(-dCA/dt)

params, cov = np.polyfit(x, y, 1, cov=True)

a = params[0]
b = params[1]

print(f"alpha = {a:.3f}")
print(f"ln k = {b:.3f}")
print(f"k = {np.exp(b):.3f}")

x_values = np.linspace(0, x[-1], 100)
y_values = x_values * params[0] + params[1]
plt.figure(figsize=(12, 8))
plt.plot(x, y, 'o', color='black', label="Datapunkter")
plt.plot(x_values, y_values, color='black', label="Regresjonskurve")
ax = plt.gca()
ax.tick_params(axis='both', which='major', labelsize=15)
ax.tick_params(axis='both', which='minor', labelsize=12)
ax.spines['right'].set_color('none')
ax.spines['top'].set_color('none')
plt.xlabel(r"$\ln \left(\overline{C_A}\right)$", size=23)
plt.ylabel(r"$\ln \left(-\frac{\Delta C_A}{\Delta t}\right)$", size=25)
plt.legend(fontsize=20)
plt.show()

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